

IN THE CLAIMS:

1. (Currently Amended) A method for supplying zinc ions to an alkaline zinc plating bath, wherein comprising the steps of bringing a source of zinc ions and a zinc dissolution accelerating metal are brought into electrically direct or indirect contact with each other, and shaking, vibrating, or rotating the source of zinc ions and the zinc dissolution accelerating metal ~~are shaken, vibrated or rotated~~ in a plating solution in order to accelerate zinc dissolution from the source of zinc ions.

2. (Currently Amended) The method according to claim 1, wherein comprising placing the source of zinc ions and the zinc dissolution accelerating metal are put in a same vessel, brought and bringing the source of zinc ions and the zinc dissolution accelerating metal into direct contact with each other and moved moving in the vessel in the plating solution by shaking, vibrating or rotating the vessel in order to accelerate zinc dissolution.

3. (Currently Amended) The method according to claim 1, wherein comprising placing the source of zinc ions and the zinc dissolution accelerating metal are put in different vessels, connected via an electric conductor, and moved moving the source of zinc ions and the zinc dissolution accelerating metal in the vessels in the plating solution by shaking, vibrating or rotating the vessels in order to accelerate zinc dissolution.

4. (Currently Amended) The method according to ~~any of the claims~~ claim 1 to 3, wherein the source of zinc ions comprises at least one or more kinds selected from of zinc, and zinc alloy, zinc or zinc alloy whose surface is plated or contacted with metal that is more electropositive than zinc, and the zinc dissolution accelerating metal comprises[;] 1) metal that is more electropositive than zinc, 2) metal that is more electropositive than zinc, with which one or more kinds selected from of iron, cobalt, nickel, carbon, silicon, manganese,

chromium, molybdenum and tungsten are contacted, 3) metal that is more electropositive than zinc, in which one or more ~~kinds selected from~~ of iron, cobalt, nickel, carbon, silicon, manganese, chromium, molybdenum and tungsten are dispersed, 4) alloy comprising metal that is more electropositive than zinc and one or more ~~kinds selected from~~ of iron, cobalt, nickel, carbon, silicon, manganese, chromium, molybdenum and tungsten, or 5) a composite or mixture of two or more ~~kinds~~ members selected from above 1) to 4).

5. (Currently Amended) The method according to ~~any of the claims~~ claim 1 to 3, wherein comprising controlling a degree of contact ~~both or either of~~ at least one of the source of zinc ions and the zinc dissolution accelerating metal ~~have with a~~ the plating solution ~~is controlled~~ in response to an analysis of a zinc ion concentration in the plating solution to adjust the zinc ion concentration in the plating solution.

6. (Currently Amended) The method according to claim 4, wherein comprising controlling the degree of contact ~~both or either of~~ at least one of the source of zinc ions and the zinc dissolution accelerating metal ~~have with the~~ plating solution ~~is controlled~~ in response to the analysis of the zinc ion concentration in the plating solution to adjust the zinc ion concentration in the plating solution.

7. (New) The method of claim 4, wherein the source of zinc ions comprise at least one of zinc and zinc alloy whose surface is plated or contacted with metal that is more electropositive than zinc.